

<p>Lithium (Li)</p> <p><u>atomic weight = 7</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 59</i></p>	<p>Beryllium (Be)</p> <p><u>atomic weight = 9</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: -241</i></p>	<p>Boron (B)</p> <p><u>atomic weight = 11</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 15</i></p>	<p>Carbon (C)</p> <p><u>atomic weight = 12</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 122</i></p>
<p>Nitrogen (N)</p> <p><u>atomic weight = 14</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: -209</i></p>	<p>Oxygen (O)</p> <p><u>atomic weight = 16</u></p> <p>Oxygen Combination: <b>1:3</b></p> <p><i>Property:</i> <i>Electron Affinity: 141</i></p>	<p>Fluorine (F)</p> <p><u>atomic weight = 19</u></p> <p>Oxygen Combination: <b>2:7</b></p> <p><i>Property:</i> <i>Electron Affinity: 330</i></p>	<p>Sodium (Na)</p> <p><u>atomic weight = 23</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 52</i></p>
<p>Magnesium (Mg)</p> <p><u>atomic weight = 24</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: -232</i></p>	<p>Aluminum (Al)</p> <p><u>atomic weight = 27</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 120</i></p>	<p>Silicon (Si)</p> <p><u>atomic weight = 28</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 133</i></p>	<p>Phosphorus (P)</p> <p><u>atomic weight = 31</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: 44</i></p>
<p>Sulfur (S)</p> <p><u>atomic weight = 32</u></p> <p>Oxygen Combination: <b>1:3</b></p> <p><i>Property:</i> <i>Electron Affinity: 200</i></p>	<p>Chlorine (Cl)</p> <p><u>atomic weight = 36</u></p> <p>Oxygen Combination: <b>2:7</b></p> <p><i>Property:</i> <i>Electron Affinity: 350</i></p>	<p>Potassium (K)</p> <p><u>atomic weight = 39</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 48</i></p>	<p>Calcium (Ca)</p> <p><u>atomic weight = 40</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: -156</i></p>
<p>Gold (Au)</p> <p><u>atomic weight = 197</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 222</i></p>	<p>Titanium (Ti)</p> <p><u>atomic weight = 48</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 19</i></p>	<p>Vanadium (V)</p> <p><u>atomic weight = 51</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: 1890</i></p>	<p>Chromium (Cr)</p> <p><u>atomic weight = 52</u></p> <p>Oxygen Combination: <b>1:3</b></p> <p><i>Property:</i> <i>Electron Affinity: 64</i></p>

<p>Manganese (Mn)</p> <p><u>atomic weight = 55</u></p> <p>Oxygen Combination: <b>2:7</b></p> <p><i>Property:</i> <i>Electron Affinity: 0</i></p>	<p>Copper (Cu)</p> <p><u>atomic weight = 63</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 118</i></p>	<p>Zinc (Zn)</p> <p><u>atomic weight = 65</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 0</i></p>	<p>Mercury (Hg)</p> <p><u>atomic weight = 200</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 0</i></p>
<p>Thallium (Tl)</p> <p><u>atomic weight = 204</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 430</i></p>	<p>Arsenic (As)</p> <p><u>atomic weight = 75</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: 817</i></p>	<p>Selenium (Se)</p> <p><u>atomic weight = 78</u></p> <p>Oxygen Combination: <b>1:3</b></p> <p><i>Property:</i> <i>Electron Affinity: 194</i></p>	<p>Bromine (Br)</p> <p><u>atomic weight = 80</u></p> <p>Oxygen Combination: <b>2:7</b></p> <p><i>Property:</i> <i>Electron Affinity: 324</i></p>
<p>Rubidium (Rb)</p> <p><u>atomic weight = 85</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 46</i></p>	<p>Strontium (Sr)</p> <p><u>atomic weight = 87</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: -167</i></p>	<p>Yttrium (Y)</p> <p><u>atomic weight = 88</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 1500</i></p>	<p>Zirconium (Zr)</p> <p><u>atomic weight = 90</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 48</i></p>
<p>Niobium (Nb)</p> <p><u>atomic weight = 94</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: 2500</i></p>	<p>Molybdenum (Mo)</p> <p><u>atomic weight = 96</u></p> <p>Oxygen Combination: <b>1:3</b></p> <p><i>Property:</i> <i>Electron Affinity: 97</i></p>	<p>Lead (Pb)</p> <p><u>atomic weight = 207</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 106</i></p>	<p>Silver (Ag)</p> <p><u>atomic weight = 108</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 125</i></p>
<p>Cadmium (Cd)</p> <p><u>atomic weight = 112</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 0</i></p>	<p>Indium (In)</p> <p><u>atomic weight = 114</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 500</i></p>	<p>Tin (Sn)</p> <p><u>atomic weight = 118</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 121</i></p>	<p>Antimony (Sb)</p> <p><u>atomic weight = 122</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: 630</i></p>

<p>Tellurium (Te)</p> <p><u>atomic weight = 128</u></p> <p>Oxygen Combination: <b>1:3</b></p> <p><i>Property:</i> <i>Electron Affinity: 190</i></p>	<p>Iodine (I)</p> <p><u>atomic weight = 127</u></p> <p>Oxygen Combination: <b>2:7</b></p> <p><i>Property:</i> <i>Electron Affinity: 300</i></p>	<p>Cesium (Cs)</p> <p><u>atomic weight = 133</u></p> <p>Oxygen Combination: <b>2:1</b></p> <p><i>Property:</i> <i>Electron Affinity: 45</i></p>	<p>Barium (Ba)</p> <p><u>atomic weight = 137</u></p> <p>Oxygen Combination: <b>1:1</b></p> <p><i>Property:</i> <i>Electron Affinity: -52</i></p>
<p>Lanthanum (La)</p> <p><u>atomic weight = 138</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 1800</i></p>	<p>Cerium (Ce)</p> <p><u>atomic weight = 140</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 0</i></p>	<p>Tantalum (Ta)</p> <p><u>atomic weight = 182</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: 3000</i></p>	<p>Tungsten (W)</p> <p><u>atomic weight = 184</u></p> <p>Oxygen Combination: <b>1:3</b></p> <p><i>Property:</i> <i>Electron Affinity: 58</i></p>
<p>Bismuth (Bi)</p> <p><u>atomic weight = 208</u></p> <p>Oxygen Combination: <b>2:5</b></p> <p><i>Property:</i> <i>Melting Point: 271</i></p>			
<p>*Scandium (Sc)</p> <p><u>atomic weight = 44</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 1000</i></p>	<p>*Gallium (Ga)</p> <p><u>atomic weight = 68</u></p> <p>Oxygen Combination: <b>2:3</b></p> <p><i>Property:</i> <i>Boiling Pt of Salt: 200</i></p>	<p>*Germanium (Ge)</p> <p><u>atomic weight = 72</u></p> <p>Oxygen Combination: <b>1:2</b></p> <p><i>Property:</i> <i>Electron Affinity: 116</i></p>	<p>*Technetium (Tc)</p> <p><u>atomic weight = 100</u></p> <p>Oxygen Combination: <b>2:7</b></p> <p><i>Property:</i> <i>Electron Affinity: 68</i></p>

Note to organizer:

If you want to reuse this activity, photocopy on colored cardstock. The final cards (with asterisk, \*) should be a different color than the rest of the cards for ease of finding and retrieving before giving an envelope of cards to the students. For example: make copies on blue, pink, green, purple cardstock. The team who gets the pink set would get the green \* cards, the team with the purple set would get the blue \* cards, etc. Keep the \* cards in a separate envelope and give to students after they have made their predictions in part 1.